
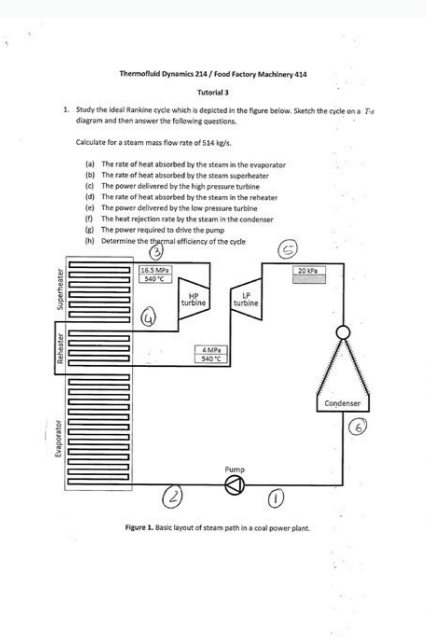


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1. Let's consider the equation



Write the molar ratios for

- a CH₄/ H₂O
- b H₂O/ H₂
- c H₂/ CH₄

- I. If you use 1 mole of CH₄ how many moles of H₂ could be produced?
- II. If 7 moles of H₂ were produced, how many moles of H₂O would be required?
- III. If 62.0g of CO were produced, how many moles of CH₄ would have been required?

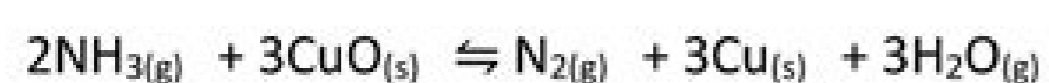
2. Let's consider the equation



Write the molar ratios for

- a. H₂/S₈
- b. H₂/H₂S
- c. S₈/H₂S
- I. If you had 3.5 moles of H₂ available and excess S₈ what is the most number of moles of H₂S you could produce?
- II. What would be the mass of this maximum yield?
- III. If you wanted to produce 3.00kg of H₂S what mass of S₈ would be required?

3. Let's consider the equation



Write the molar ratios for

- a. NH₃/ CuO
- b. CuO/ NH₃



Introduction to Statistics – Chapter 1

Applied Research Methods in Social Sciences – Faculty of Political Sciences - University Babeş-Bolyai Cluj

Beca Maruša

Answer ex 12

Independent variable = oxytocin

Dependent variable = trust

13. For each of the following, determine whether the variable being measured is discrete or continuous and explain your answer.

- Social networking (number of daily minutes on Facebook)
- Family size (number of siblings)
- Preference between digital and analog watch
- Number of correct answers on a statistics quiz

Answer ex 13

- Social networking is a **continuous variable** because it can measure fractions of minutes like seconds.
- Family size is a **discrete variable** because the number of siblings can't take a value between integer numbers, for example between 3 and 4.
- Preference between digital and analog watch is a **discrete** variable because it consists of separate, indivisible categories.
- Number of correct answers on a statistics quiz is a **discrete** variable because there are a finite and countable number of questions.

14. Four scales of measurement were introduced in this chapter: nominal, ordinal, interval, and ratio.

- What additional information is obtained from measurements on an ordinal scale compared to measurements on a nominal scale?
- What additional information is obtained from measurements on an interval scale compared to measurements on an ordinal scale?
- What additional information is obtained from measurements on a ratio scale compared to measurements on an interval scale?

Answer ex 14

- An **ordinal scale** consists of a set of categories that are organized in an ordered sequence. Measurements on an ordinal scale rank observations in terms of size or magnitude. The categories that make up an **ordinal scale** not only have different names (as in a **nominal scale**) but also are organized in a fixed order corresponding to differences of magnitude. With measurements from an ordinal scale, you can determine whether two individuals are different and you can determine the direction of difference.
- An **interval scale** consists of ordered categories that are all intervals of exactly the same size. Equal differences between numbers on the scale reflect equal differences in magnitude.

Page 6 of 11

CHEMISTRY XI NOTES INTRODUCTION TO FUNDAMENTAL CONCEPTS OF CHEMISTRY - QUESTIONS ANSWERS

Chapter # 01 Chemistry - XI

Questions Answers

INTRODUCTION TO FUNDAMENTAL CONCEPTS OF CHEMISTRY

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Q.1: Give reasons that formation of positive ion is endothermic reaction?

Ans: Formation of positive ion is an endothermic process because heat is required to remove electron from a neutral atom to form positive ion.

Q.2: Why a physical change is temporary change?

Ans: Because due to the physical change no new substances are formed, therefore it is temporary change.

Q.3: Explain why chemical change is permanent change?

Ans: As during chemical change new substances are formed so it is a permanent change.
Example:
➤ Rusting of iron.
➤ Conversion of water in a steam.

Q.4: Why a limiting reactant controls the quantities of products in a chemical reaction?

Ans: The limiting reactant is a substance which produces least of products after its complete consumption. After the complete consumption of the limiting reactant no more products are formed, as the reactants are not mixed in stoichiometric amounts to form products. During a chemical reaction, thus a limiting reactant controls the quantities of products in a chemical reaction.

Q.5: What is the atomic mass unit?

Ans: It is a unit of mass used for atoms and molecules and it equal to the 12 of the mass of carbon atom - 12. It is obtained by dividing the unity by Avogadro's number (6.02×10^{23}).

Q.6: Define empirical formula and explain how it is related to molecular formula?

Ans: The simplest formula of a substance which tells us the ratio of atoms in the molecule of a substance is called empirical formula.
 $MF = n(E.F)$

practicalcentre.blogspot.com Page 11

What is the introduction to chemistry. What are introductory questions.

CH 16-17. Chemical bond; Structure and shape. * Worksheet on ions and ionic compounds. The following connections are to all the main dispensations Intro Chem (X11) for the entire semester; They are in chronological order. Atomic theory: the quantum model of the atom. Web page: Designation of chemicals, with answers. The temperature in which a substance changes from a solid to a liquid. The dispenser can contain information on homework, priority material within the chapter, known errors and other comments. Most of these links are "Handout Chapters" and are here as a PDF file. For further questions about the practice, consult the quizzes of the practice. For all intro chem tests you can - and they should use a periodic table. CH 20. * The periodic table has a table is the one published here on my download page. CH 11-12. Density, volume and weight A defined shape and a volume A defined volume, but not a defined form a defined form, but not a volume defined a shape and a variable volume A Which of the following is a list of compounds? CH 18-19. Set # 1, autumn 2005 test or response RTF button RTF file test cover (chapter titles, with cracolozio 2 / E book key); Other notes. Chemical element Chemical reaction organic chemistry chemistry basic chemistry Which of these passages of the scientific method implies to affirm a relationship but not proposing an explanation for the relationship? The sample tests are available in the upper section of this page. CH 1-2. The accompanying graph shows the data. * Sample test: I distribute the sample for testing 1; This test and your response key, all other tests and sample keys and other test information, are in the section above: sample tests, with response keys and other test information (Chem). Test (and keys) are published here in RTF. The approximate coverage of each test is reported below. Chemical formula relationships: Reactions and equations: Quantity reports in chemical reactions; Atomic theory: the quantum quantum model The atom. * Test includes the "lid", which explains the procedure for a test to domicile. (Annex A CH 6 Handout.) The response keys for the "Nominating Chemicals" worksheet are published. Intro Chem (X11) Home Contact Information A, A, A, A, Site Home page Last Update: June 26, 2019 Chemistry is the "Scientific Study of Matter, Its Properties, and Interactions with Other Matter and Energy." Chemistry is the scientific study of matter and its components. * Designation of chemicals. Some "test information" that delivers together with the first sample test: test information. See the "Handout of the Chapter" section of the course program for more information on these lizards. Update this answer after claiming an answer that I would have 24 hours to send a draft. This is presumably portable and should be ok for users of a range of word processors. (Introduction to chemistry; matter and energy, measurement and chemical calculations; Atomic theory: The nuclear model of the atom; Chemical nomenclature.) Response button for test # 1 Key 1 RTF test # 2 test 2 RTF * CH 6-10. (Chemical nomenclature. Response button for test # 2 Button 2 RTF test N. 3 (final exam) Test 3 RTF CH 11, 12, 4, 13, 15-19. * This dispensation also includes a set of additional questions for CH 16, which are in a separate file: CH 16 Additional questions. Quantitative reports in chemical reactions. The worksheet and the graph will open each in a new window as a PDF file. (My contact information is at the bottom of each of each of the My web pages.) Always help you understand how to answer to the question and why. Acid-base reactions (proton transfer); oxidation reduction reactions (redox). (A spreadsheet includes the graph, in addition to the table on which is based.) CH 8. (Introduction to matter and energy; measurement and chemical calculations; Atomic theory: the nuclear model of the atom; Chemical nomenclature.) * This test will be delivered even in the classroom. Other course course Like quiz practices and additional discussions of various topics, they are available at the introduction home page. (For class use, they are connected to CH 7 Handout.) Response card for the crust worksheet. CH 9. Introduction to gases; Ideal gas and applications; Gas, liquids and solids. Test A, A, sample tests, with the response keys and other test information (Intro Chem) The sample tests are real tests of previous classes. The purpose of doing it is so that you can see not only the content of the test, but also what looks like a real test.) Attention: exactly what it is covered, or even the order of the topics, it is not always D Agreement from one semester another. Atomic theory; the nuclear model of the atom. Rating: 0% Rank: Correct Rank: Correct Response: Ratings: 0% Rank: Correct Response: Test (sample test, with the response keys and other test information) Heats: Various A e lower part of the page; Return the links and contact information Those who use my website materials for self-studio ... two complete sets are published below, with response keys. * CH 9 Handout also includes a set of additional questions, which are in a separate file: CH 9 Additional questions. Test n. 1 test 1 RTF * CH 1, 2, 3, 5 and part of 6. You are welcome - and encouraged - to ask me questions when difficulties arise. Furthermore, it gives me a feeling for the level in which you are facing the question - which can vary depending on your level and level of the course. CH 5. Complete set complete set Zip RTFS; Contains all the files listed below. (I will also distribute a sample for test no. 1 to make sure everyone can see at least the first before entering the test. Information on the note page, which you can use on all tests. An Editor The presentation and publish your submission or provide feedback. Please let me know if you have problems accessing any file; Sometimes, your comment is the way you discover it discover There is a problem. CH 4-13-15. Chemical balance. * Worksheet: explore the chemistry of the terrestrial crust. * Using your scientific calculator. Therefore, you should not rely on the sample tests as precise guides to test the content. Introduction to chemistry: introduction to active learning; Matter and energy. Solutions; Net ionic equations. Next Response Chapter 1 - The Chemical World - Exercises - Questions - Page 10; 6 Answer Previous Chapter 1 - The Chemical World - Exercises - Questions - Page 10; 4 Score: 0% Rank: Correct Response: 10 Questions | Total attempts: 3860 Which of the following is not part of the scientific method? Much of the content is general and would be useful with other editions and even other books. Water, carbon dioxide, hydrogen benzene, nitrogen, oxygen the temperature in which a substance changes from a liquid to a gas all that has a defined volume but no defined shape density is a measure of how many masses is contained in a given Volume of unity The quantity volume of space something takes into consideration. (Annex A CH 6 Handout.) CH 7. For more information on the editions of the Cracolice book, refer to the textbook information on the Informative Chemical Introductory (X11) information page. Reactions and equations. Chemical formula reports. In addition to individual files, I published a zip archive file containing the entire test set and keys. You can help us review, improving and updating this answer. CH 10. CH 3. CH 6. Top of Page The Practical Chemical Program Problems Link link to all the sets of chimo problems that are rejected to my website. (Chemical bond; structure and shape; gas introduction; the ideal gas law and its applications; gases, liquids and solids; solutions; net ionic equations; acid-base reactions (proton-transfer); oxidation reduction reactions (redox).) Key Answer for test no. 3 key 3 set RTF # 2, spring test 2005 or response button RTF file test cover (chapter titles, with key wrench 2 / and book); Other notes. Response button for test # 1 key 1 RTF test No. 2 test 2 RTF * CH 6-10. In fact, the first page includes an expanded introduction and includes questions. If you find formatting problems with these files, let me know. (Chemical bond; structure and shape; gas introduction; the ideal gas law and its applications; gases, liquids and solids; solutions; net ionic equations; acid-base reactions (proton-transfer); oxidation reduction reactions (redox).) Response key for test # 3 Key 3 RTF A, A, a, top of page heats; Miscellaneous (Intro Chem) For each chapter provided a short pantry that offers my comments on the chapter. Here is what students for tests: periodic table. These dispenses are specifically implored to the second edition of the Cracolice book. * The dispenser on the dimensional analysis (a supplement for CH 3) is published on the page of chemistry practice problems, along with some other self-help materials. This allows me to answer what you are thinking, it allows me to concentrate my answer to where you are having problems. Chemical measurement and calculations. Chemical nomenclature. The level of speech - and your learning of chemistry - has improved trying to focus on reasons, not simply answers. Test n. 1 test 1 RTF CH 1, 2, 3, 5 and part of 6. 6.

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